

	Cl ⁻	CO ₃ ⁻²	OH ⁻	SO ₄ ⁻²	PO ₄ ⁻³	NO ₃ ⁻	P ⁻³
Fe ³⁺	FeCl ₃	Fe ₂ (CO ₃) ₃	Fe(OH) ₃	Fe ₂ (SO ₄) ₃ FePO ₄	Fe(NO ₃) ₃	FeP	
	Iron (III) Chloride		Iron (III) Hydroxide				Iron (III) Phosphide
Al ³⁺							
Co ³⁺							
Fe ²⁺	FeCl ₂	FeCO ₃	Fe(OH) ₂	FeSO ₄	Fe ₃ (PO ₄) ₂ Fe(NO ₃) ₂	Fe ₃ P ₂	
	Iron (II) Chloride						
H ⁺							
		Careful ↓					
Au ⁴⁺	AuCl ₄	Au(CO ₃) ₂	Au(OH) ₄	Au(SO ₄) ₂	Au ₃ (PO ₄) ₄	Au ₃ (P ₂) ₄ or AuP ₄	
	Gold (III) Chloride						

Gold (IV)
hydroxide

Gold (IV)
Sulfate

Gold (III)
Phosphide

Standard: #3-1

WRITING IONIC FORMULAS

Include the name and formula of each chemical

	Cl ⁻ Chloride	CO ₃ ⁻²	OH ⁻	SO ₄ ⁻²	PO ₄ ⁻³	NO ₃ ⁻	P ⁻³
Na ⁺	NaCl	Na ₂ CO ₃	NaOH	Na ₂ SO ₄	Na ₃ PO ₄	NaNO ₃	Na ₃ P
	Sodium Chloride	Sodium Carbonate	Sodium hydroxide	Sodium Sulfate	Sodium Phosphate	Sodium Nitrate	Sodium Phosphide
NH ₄ ⁺ Ammonia	NH ₄ Cl	(NH ₄) ₂ CO ₃	NH ₄ OH	(NH ₄) ₂ SO ₄	(NH ₄) ₂ PO ₄	NH ₄ NO ₃	(NH ₄) ₃ P
	Ammonium Chloride	Ammonium Carbonate	Ammonium hydroxide	Ammonium Sulfate	Ammonium Phosphate	Ammonium Nitrate	Ammonium Phosphide
K ⁺							
	<p>Notice careful common errors</p> <p>↓ ↓</p>						
Ca ²⁺	CaCl ₂	CaCO ₃	Ca(OH) ₂	CaSO ₄	Ca ₃ (PO ₄) ₂	Ca(NO ₃) ₂	Ca ₃ P ₂
Mg ²⁺			Mg(OH) ₂				
Zn ²⁺			Zn(OH) ₂				